

PATENT ABSTRACTS OF JAPAN

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CLAIMS

[Claim(s)]

[Claim 1] The secondary battery that use a positive active material expressed by $\text{Li}_a\text{Ni}_b\text{M}^1_c\text{M}^2_d\text{M}^3_e\text{O}_2$ (wherein, M^1 is at least one sort of elements chosen out of Co, Mn, and Fe, M^2 is at least one sorts of elements chosen out of B, Al, In, and Sn, and M^3 is at least one sort of elements chosen out of Mg, Zn.

DETAILED DESCRIPTION

[Detailed Description of the Invention]

[0009] [The purpose of this invention]

To offer the long lasting lithium secondary battery which has an excellent rate capability.

[0021] (Example 1- coin cell A1)

- $\text{LiOH} \cdot \text{H}_2\text{O}$, Ni_2CO_3 , CoCO_3 , B_2O_3 , and MgO were used.
- The mole ratio, $\text{Li} : \text{Ni} : \text{Co} : \text{B} : \text{Mg}$ is $1.03 : 0.88 : 0.10 : 0.01 : 0.01$.
- The mixture was calcinated at 750 degrees C under oxygen atmosphere for 20 hours. After calcination, it was cooled in dry air, and was milled.
- Cathode : Acetylene black : PTFE = 85 : 10 : 5
Positive electrode is dried at 200 degree C in vacuum.
Counter electrode is Li metal.
Electrolyte is EC/DEC (1:1 vol%)

A02

[0024] (Example 2 - coin cell A2)

- $\text{LiOH} \cdot \text{H}_2\text{O}$, Ni_2CO_3 , CoCO_3 , $\text{Al}_2(\text{NO}_3)_3$, and MgO were used.
- The mole ratio, $\text{Li} : \text{Ni} : \text{Co} : \text{Al} : \text{Mg}$ is $1.03 : 0.88 : 0.10 : 0.01 : 0.01$.

[0025] (Example 3 - coin cell A3)

- $\text{LiOH} \cdot \text{H}_2\text{O}$, Ni_2CO_3 , CoCO_3 , $\text{In}(\text{NO}_3)_3 \cdot x\text{H}_2\text{O}$ and MgO were used.
- The mole ratio, $\text{Li} : \text{Ni} : \text{Co} : \text{In} : \text{Mg}$ is $1.03 : 0.88 : 0.10 : 0.01 : 0.01$.

[0026] (Example 4 - coin cell A4)

- $\text{LiOH} \cdot \text{H}_2\text{O}$, Ni_2CO_3 , CoCO_3 , SnO and MgO were used.
- The mole ratio, $\text{Li} : \text{Ni} : \text{Co} : \text{Sn} : \text{Mg}$ is $1.03 : 0.88 : 0.10 : 0.01 : 0.01$.

[0027] (Example 5 - coin cell A5)

- $\text{LiOH} \cdot \text{H}_2\text{O}$, Ni_2CO_3 , CoCO_3 , B_2O_3 , and ZnO were used.
- The mole ratio, $\text{Li} : \text{Ni} : \text{Co} : \text{B} : \text{Zn}$ is $1.03 : 0.88 : 0.10 : 0.01 : 0.01$.

[0028] (Example 1 of a comparison – coin cell B1)

- $\text{LiOH} \cdot \text{H}_2\text{O}$, Ni_2CO_3 were used.
- The mole ratio, $\text{Li} : \text{Ni}$ is $1.03 : 1.00$.

[0029] (Example 2 of a comparison – coin cell B2)

- $\text{LiOH} \cdot \text{H}_2\text{O}$, Ni_2CO_3 , CoCO_3 were used.
- The mole ratio, $\text{Li} : \text{Ni} : \text{Co}$ is $1.03 : 0.90 : 0.10$.

[0030] (Example 3 of a comparison – coin cell B3)

- $\text{LiOH} \cdot \text{H}_2\text{O}$, Ni_2CO_3 , B_2O_3 were used.
- The mole ratio, $\text{Li} : \text{Ni} : \text{B}$ is $1.03 : 0.90 : 0.10$.

[0031] (Example 4 of a comparison – coin cell B4)

- $\text{LiOH} \cdot \text{H}_2\text{O}$, Ni_2CO_3 , CoCO_3 , B_2O_3 were used.
- The mole ratio, $\text{Li} : \text{Ni} : \text{Co} : \text{B}$ is $1.03 : 0.89 : 0.10 : 0.01$

[0034]

[Table 1]

電池	放電電流 3mA 放電容量(mAh)		放電電流 10mA 放電容量(mAh)	
	1st cycle	10th cycle	1st cycle	10th cycle
A1	65	64	58	55
A2	66	63	57	54
A3	65	64	57	55
A4	66	64	56	53
A5	66	63	55	52
B1	52	45	28	25
B2	50	43	29	24
B3	49	40	30	25
B4	63	62	30	28

Discharge current = 3mA
Discharge capacity (mAh)
1st cycle 10th cycle

Discharge current = 10mA
Discharge capacity (mAh)
1st cycle 10th cycle

[0035] As shown in Table 1, the cells A1, A2, A3, A4 and A5 by this invention had a large initial discharge capacity compared with the comparison cell B1, B2, and B3. Furthermore, the cells A1, A2, A3, A4 and A5 by this invention had a good rate capability compared with the cell B4.